

Classifying Matter

Instructions for the Element, Compound, and Mixture Foldable

1. Neatly cut out the 3-tab foldable
 2. Cut in between lines to make 3 tabs
 3. Glue **ONLY** on the title line and attach the 3-tab foldable to your I.N.B.
 4. Follow the directions carefully
 5. Use your textbook to help write the bulleted information.
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1. Label each section's **Cover** in this order: **Elements, Compounds, Mixtures**
 2. **Element Cover (color)**
 - ⇒ draw and label 2 examples of elements
 - ⇒ write definition of element below pictures
 3. **Compound Cover (color)**
 - ⇒ draw Figure 2 (pg. 61) include labels (Na Cl and NaCl)
 - ⇒ write definition of compound below your drawing
 4. **Mixtures Cover (color)**
 - ⇒ Draw Chocolate Chip Cookie
 - ⇒ Draw salt-water solution (looks same throughout –label solution)
 - ⇒ Write definition of heterogeneous and homogenous mixture under pictures (Check the board for definition helps)
 5. **Element Inside (write the following in bullets)**
 - ⇒ An element cannot be broken down into simpler substances by physical or chemical means
 - ⇒ Each element can be classified by a unique set of physical and chemical characteristics.
 - ⇒ Based on their properties elements are divided into metals, nonmetals, and metalloids
 6. **Compound Inside (write the following in bullets)**
 - ⇒ Compounds can only be separated by chemical means such as heat, electrolysis, or a reaction.
 - ⇒ Compounds have a unique set of physical and chemical properties that differ from the elements that make up the compound.
 - ⇒ Elements join in specific ratios according to their masses to form compounds.
 7. **Mixtures Inside (write the following in bullets)**
 - ⇒ A Mixture can be separated by physical means such as sorting, filtration, and temperature.
 - ⇒ A mixture has the same properties of the substances that make up the mixture.
 - ⇒ A homogenous mixture is a mixture that appears to be the same throughout.
 - ⇒ A heterogeneous mixture may have a different composition in every sample.